

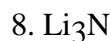
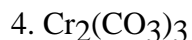
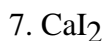
## Chapter 9 Practice #2

Name \_\_\_\_\_

Period \_\_\_\_\_ Date \_\_\_\_\_

	Name of Cation	Name of Anion	Formula of Cation	Formula of Anion	Formula of Compound	Name of Compound
1	Calcium ion	Chloride ion				
2	Iron(III) ion	Phosphide ion				
3			$\text{Na}^{1+}$	$\text{S}^{2-}$		
4			$\text{Al}^{3+}$	$\text{Br}^{1-}$		
5						Lithium Sulfide
6						Platinum (IV) Oxide
7	Magnesium ion	Carbonate ion				
8			$\text{Ca}^{2+}$	$\text{NO}_3^{1-}$		
9					HgS	
10					$\text{Th}_3(\text{PO}_4)_2$	

Write the names of these ionic compounds



Write the formulas for these ionic compounds

11. Chromium (IV) oxide

13. Aluminum oxide

15. Silver (I) sulfide

17. Nickel (II) sulfate

19. Potassium dichromate

12. Chromium (III) oxide

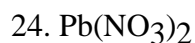
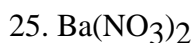
14. Nickel (II) bromide

16. Magnesium chloride

18. Iron (III) phosphate

20. Lead (IV) hydroxide

Write the names of these ionic compounds



Write the formulas of these ionic compounds.

31. Ruthenium (III) nitride

33. Calcium nitrite

35. Manganese (II) sulfite

37. Cobalt (III) chromate

39. Rubidium cyanide

32. Titanium (IV) nitrate

34. Sodium sulfate

36. Rhenium (II) phosphate

38. Yttrium (II) dichromate

40. Nickel (II) acetate